Roles of structural defects in polycrystalline platinum nanowires for enhanced Oxygen Reduction Activity

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Abstract

Heterogeneous nanocatalysts have uneven surface reactivity and lively change in reactions, making the disclosure of their catalytic origins highly challenging. Phenomenologically, Pt-based nanowires, one of the most important nanostructures, have been noted to display high activity for the oxygen reduction reaction (ORR); however, the fundamentally catalytic origin behind such high performance remains elusive and was implicitly ascribed to one-dimensional structure and/or low-index facets. Here, with molecular-level and operando spectroscopic evidence, we reveal that the polycrystalline platinum nanowires with enhanced ORR activity are closely correlated with structural defects including grain boundaries (GBs) and atomic steps, which is previously neglected or simplified. The local lattice strain is induced by GBs and increases with the misorientation angles of GBs. Despite the same coordination number between edge sites and atomic steps, the latter disturbs interfacial water networks via preferable hydrogen bonding, which destabilizes ORR intermediates on terraces and presumably promotes proton transfer.

Keywords: oxygen reduction reaction, polycrystalline nanowires, atomic steps, grain boundaries, catalytic defects

1. Introduction

Electrocatalysts are crucial in future energy conversion and storage techniques. In response, rational design of efficient electrocatalysts is indispensable and heavily relies on the knowledge on molecular-level electroactive factors and their operating mechanisms. However, practical nanoelectrocatalysts are structurally heterogeneous at multi-scales, comprise diverse defects and evolve dynamically in reactions, raising forbidden challenges for an in-depth understanding of electrocatalysts. The Pt-based nanocatalysts remain most-efficient for the oxygen reduction reaction (ORR) [1-8] , a crucial reaction for fuel cells and metal-air batteries [1-22]. Typically, polycrystalline Pt-based nanowires (PPNWs), one of the most important nanostructures, have been noted to display enhanced ORR activity; however, it is unclear for underlying reasons for activity enhancement. Two general clues, low-index nanofacets[9, 16, 23] and one-dimensional (1D) nanostructure,[11, 13] have been proposed. However, some puzzles remain for a long time. At first, PPNWs display an unusual size effect. A much higher ORR activity on 1 nm Pt NWs was reported than on 4.5 nm Pt NWs[9] and 200 nm Pt nanotubes,[11, 12] distinct from those nanoparticulate catalysts.[24, 25]

More importantly, the roles of defects, like atomic steps and grain boundaries (GBs), are elusive in PPNWs. Reports showed both defect-free PPNWs with Pt (111) facet[11-13] and jagged PPNWs with strained and stepped nanofacets[26] had particularly high ORR performance. Steps binding strongly to ORR intermediates were theoretically speculated inactive or low-active for ORR.[25, 27, 28] However, the acidic ORR activity on high-index facets was reported experimentally higher than on Pt(111).[1] The GBs, thin interfacial regions between grains, are inherent in polycrystals and proposed to induce catalytically active strained domains.[29, 30] Practical issues remain unanswered regarding how to tune GBs and thus optimize catalytic performance. More importantly, defects are usually reactive and easily reconstructed in reactions, thereby calling for a site-specific and operando understanding of their electrocatalytic behaviors. For example, evidence emerged that

surface charging, for example, producing oxygenates on defective or oxophilic sites, brings about the disturbance to interfacial water networks and/or double layer[31-33] and subsequently affect intermediates adsorption[31] and/or mass transport.[32] Recent reports reveal defective or disordered $PtNi_x$ catalysts showed better ORR activity than the common $PtNi_x$ alloy at similar crystalline sizes.[29, 34-37] Despite those important advances, it remains unexplored in details for whether defects in PPNWs play positive roles, and how they operate in ORR.

Here, we explored molecular-level electroactive factors resulting in enhanced ORR activity for model PPNWs. Experimental evidence reveals PPNWs's defects comprising atomic steps and GBs benefit ORR. Atomic steps on stepped nanofacets disturb interfacial water networks, destabilizing oxygenates adsorption on terraces. The microstrain is induced by GBs and affected by misorientation angles (MOAs) of GBs, inducing higher ORR activity on more disordered/defective PPNWs. Our insights into molecular-level electroactive factors and their operation mechanisms not only rationalize the enhanced ORR activity on PPNWs but also help design of more efficiency nanoelectrocatalysts, for example, by regulating density, architectures, and interfacial chemistry of GBs.

2. Experimental

2.1. Materials and Chemicals

Platinum(II) acetylacetonate ($Pt(acac)_2$), Chloroplatinic acid hexahydrate, cetyltrimethylammonium bromide (CTAB), oleyl amine (OAm), Triton™ X-114, sodium tetrahydridoborate ($NaBH_4$) and perchloric acid (TraceSELECT) were all purchased from Sigma-Aldrich and used as received without further purification. Ultrapure water (18 MΩ cm) purified in a Millipore system was used in all experiments.

2.2. Preparation of regular PPNWs and PPNWs/C

The synthesis of regular Pt nanowires follows a previous method[14] with a slight modification. Specifically**,** 20 mg $Pt(acac)_2$, 50 mg CTAB, and 4 mL OAm were

added into 20 mL high pressure vessel. After 30 min ultrasonication of these mixtures, the reaction vessel was dipped into an oil bath pre-heated at 170 ℃. After 5 min, 20.3 mg $Mo(CO)_6$ was added into the reaction vessel to reduce $Pt(acac)_2$ at 170 ℃ for 2 h. The PPNWs product was precipitated by adding 10 mL ethanol, washed intensively with hexane, and finally dispersed in hexane.

Ketjen black carbon was dispersed in 20 ml n-butylamine solvent, to which the regular PPNWs/hexane dispersion was poured, then dispersed thoroughly via ultrasonication, and stirred at room temperature for 24 h to allow for ligand exchange between oleyl amine and n-butylamine. The product was collected by centrifugation and washed with ethanol three times. To ensure complete removal of oleyl amine and/or n-butylamine adsorbed on PPNWs, the catalyst product was heated at 200 ℃ in air for 1 h in a muffle furnace. The Pt mass percentage was estimated to be 10.0 wt% by X-ray fluorescence (XRF).

2.3. Preparation of waved PPNWs and waved PPNWs/C

The synthesis of waved Pt nanowires follows a previous method[38]. Specifically, 0.25 mL Triton™ X-114 and 1 mL H_2PtCl_6 solution (15 mg_{Pt}/mL) were added into 50 ml H_2O and intensively stirred for 5 h. Then, 20 mL of freshly prepared $NaBH_4$ solution (1.5 mg/mL) was poured into the above solution to reduce $PtCl_6^{-2}$ and initiate the growth of waved PPNWs. The reaction was continued for 10 min without stirring. The product was collected by centrifugation, washed by ethanol/water solution (V/V=1/1) and finally dispersed in ethanol/water solution.

A waved PPNWs solution was poured into 50 ml ethanol/water solution (V/V=1/1) pre-dispersed with Ketjen black carbon. The resultant mixture was stirred for 24 h to allow for the deposition of waved PPNWs on Ketjen black The catalyst was collected by centrifugation and washed by ethanol/water solution. To remove residual Triton™ X-114 completely, the catalysts product was heated in air for 2 h. The Pt mass percentage was estimated to be 15.0 wt% by XRF.

2.4. Benchmark Pt/C catalyst

A commercial Pt/C (TEC10E20E) with an average particle size of 2 nm was obtained from Tanaka Kikinzoku Kogyo (TKK) and used as a reference catalyst.

3. Results and discussion

3.1. Catalyst characterization and ORR performance

We start from two model PPNWs: regular (Figure 1a-d) and waved (Figure 1e-h) PPNWs. Regular PPNWs were synthesized through reducing $Pt(acac)_2$ by $Mo(CO)_6$ and using cetyltrimethylammonium bromide as a soft template for the PPNWs growth in oleylamine solvent[14] while waved PPNWs via reducing H_2PtCl_6 by $NaBH_4$ and using Triton™ X-114 as a template in deionized water[38] (see the details in Supporting Information (SI)). The averaged-diameters of regular and waved PPNWs are 1.3 ± 0.3 and 2.2 ± 0.6 nm, respectively (Figure S1). The most notable difference between two PPNWs is in the freedom of grains. Specifically, grains in regular PPNWs are restricted in 1D direction, while in waved PPNWs orientate quite freely in 3D directions, consequently appearing more disordered on wave PPNWs (see more in Figures S2-S3).

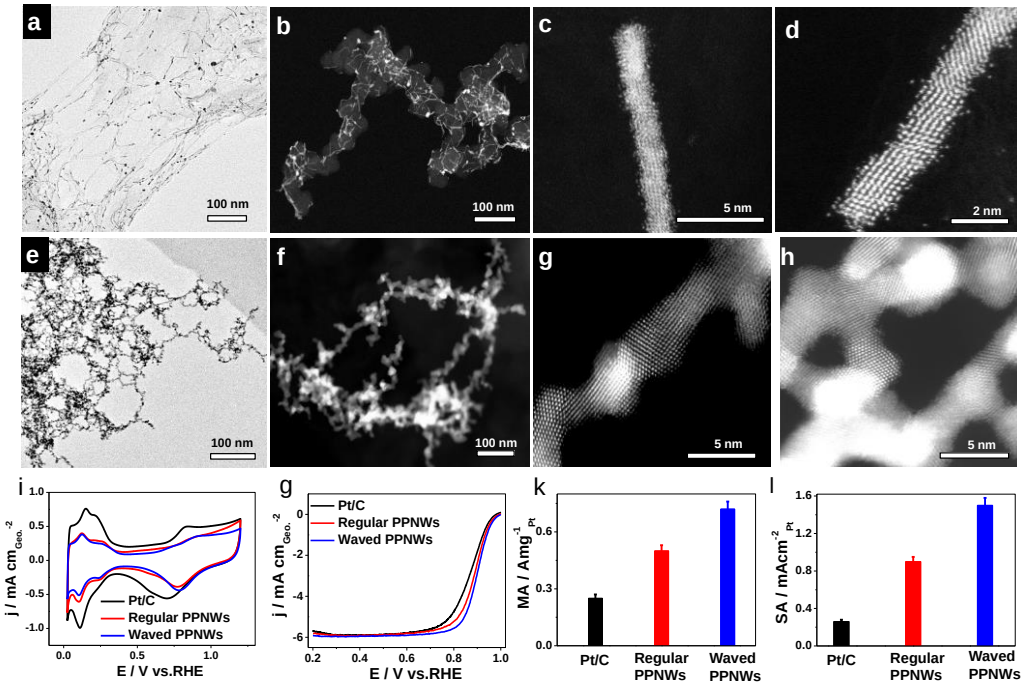

Figure 1: STEM images for (a-d) regular and (e-h) waved PPNWs, as-prepared (a, e) and carbon supported PPNWs (b-d, f-h). (i) Cyclic voltammograms measured in N_2-saturated 0.1 M $HClO_4$ aqueous electrolyte at 50 mV/s. (g) ORR anodic polarization curves measured at 20 mV/s (1600 rpm). (k-l) Histograms for mass activity (MA, k) and specific activity (SA, l) calculated by Koutecky-Levich equation to obtain mass transport-corrected kinetic current. Pt loadings on three electrodes are 12.0 μg_{Pt} cm^{-2}_{Geo}.

We tested ORR performances of carbon-supported regular and waved PPNWs and a benchmark 2.0 nm Pt/C (TEC10E20E). Both regular and waved PPNWs/C display clearly anodically-shifted potentials for *OH/*O desorption and/or surface oxide reduction against 2 nm Pt/C, seen from their cathodic scans in cyclic voltammograms in Figure 1(i). The ORR activities of regular and waved PPNWs/C surpass 2 nm Pt/C regarding Pt-based mass activity (MA, Figure 1(k)) and electrochemical areas-based specific activity (SA, Figure 1(l)). Notably, the SA activity on waved PPNWs/C is higher than on regular PPNWs/C, contrasted with the reports about an improved ORR activity on smaller-sized NWs.[9], [11, 12] Waved PPNWs with a bit larger diameter display a higher activity against regular NWs, indicating a nanostructure-induced enhancement of ORR activity rather than a size effect. Specifically, nanoscale building blocks (i.e. 1-2 nm Pt grains) are similar for

the three catalysts; the structural discrepancy accounting for their distinct ORR activities arises from grain organization patterns, i.e. isolated nanoparticles, regular PPNWs, and waved PPNWs. In other words, new structural factors and/or catalytically operating mechanisms are generated to enable enhanced ORR activity for PPNWs.

3.2. Atomistic analysis of structural defects in PPNWs

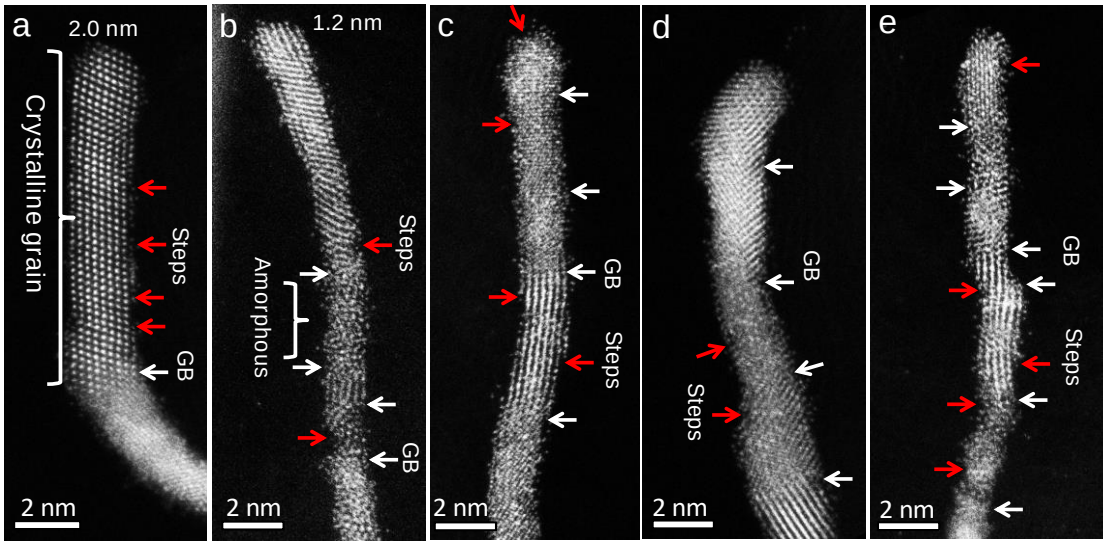

Figure 2: Atomistic observation of structural defects and heterogeneities of regular PPNWs. (a-e) HAADF-STEM images. Heterogeneous atomic structures are seen in intra- and inter-NWs, however two general defects, GBs (white arrows) and atomic steps (red arrows), are observable in all measured NWs.

An intensive atomistic observation of randomly selected PPNWs demonstrates their polycrystalline nature without any observable single-crystalline NWs (Figure 2, see more in Figure S2). Structural heterogeneity appearing intra-NWs and inter-NWs calls for a sites-specific understanding of their properties. Overall, smaller-diameter PPNWs tend to present a larger defective/disordered degree. For instance, a large domain preferably emerges at 2.0 nm PPNWs (Figure 2a), while an amorphous region arises for 1.2 nm PPNWs (Figure 2b). A thorough examination of all measured PPNWs suggests two general defects: atomic steps and GBs. Notably, most of atomic steps in PPNWs are placed in planes (denoted as the plane-step different from the edge sites of NPs in 2.0 nm Pt/C). We previously presented *in situ* spectroscopic

evidence that those plane-steps on stepped nanofacets as templates disturb interfacial water network and thus destabilize ORR intermediates on vicinal terraces.[31] By contrast, edge sites over-binding with oxygen contribute negligibly to ORR activity for Pt NPs.[25, 27, 28] As plane-steps and edge sites have same coordination number ($CN_{Pt\text{-}Pt}$=7), their distinct contributions to ORR activity presumably originate from different geometry positions. We further through in-situ spectra discuss this point below.

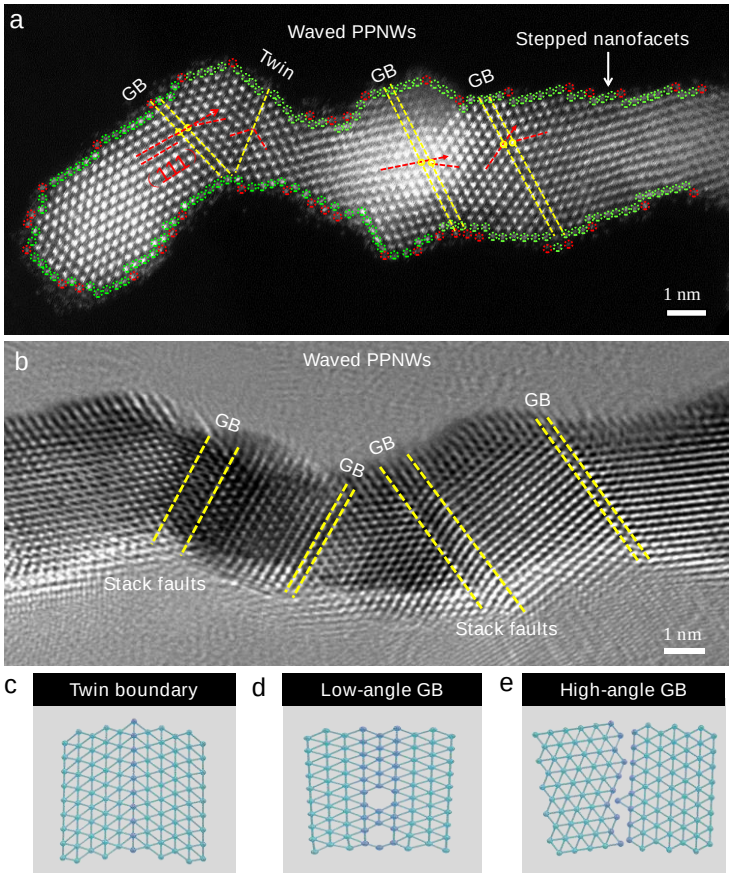

Figure 3: Atomistic observation of stepped nanofacets and GBs of waved PPNWs. (a) HAADF-STEM image and (b) high-resolution TEM image for waved PPNWs. Illustrations of structural models for (c) twin boundary, (d) low-angle GBs, and (e) high-angle GBs.[39] Red cycles in (a) suggest undercoordinated steps or kinks and green cycles represent terrace atoms.

We further measured atomistic images of waved PPNWs and found the similar results with regular PPNWs regarding plane-steps and GBs regardless of their varying nanostructures (Figures 2-3). According to the sizes of misorientation angles (MOAs)

and the excessive internal energy, GBs are classified into low-angle GBs (MOAs <15 °) and high-angle GBs (MOAs >15 °).[39] The twin is a special boundary between two grains with the relation of mirror reflection. Clearly, the grains in waved PPNWs orientate considerably freely and display large MOAs, resulting in a waved shape against regular PPNWs. Additionally, the stack fault and lattice deformation are observable around GBs (Figures 3 and 4), as GBs could create dislocation pile-ups via hindering dislocation sliding.[39, 40] Expectedly, high-angle GBs with larger MOAs via a dislocation-induced strain field cause more deformed/strained surface atoms[30] that are deviated from the ideal lattice positions and often display unusually catalytic activity. Given this, the overall larger MOAs inducing more catalytically active surface atoms may contribute to a higher ORR activity on waved PPNWs than on regular PPNWs. As the electrocatalyst structure might change dramatically in electrochemical environment and/or during electrolysis, we conducted TEM and HRTEM examination of steps and GBs in PPNWs after the ORR electrocatalysis. The TEM and HRTEM results shown in Figures S6-S7 confirm the absence of dramatic structural change for both regular and waved PPNWs and the still-existing step and GB defects in PPNWs after the ORR electrolysis.

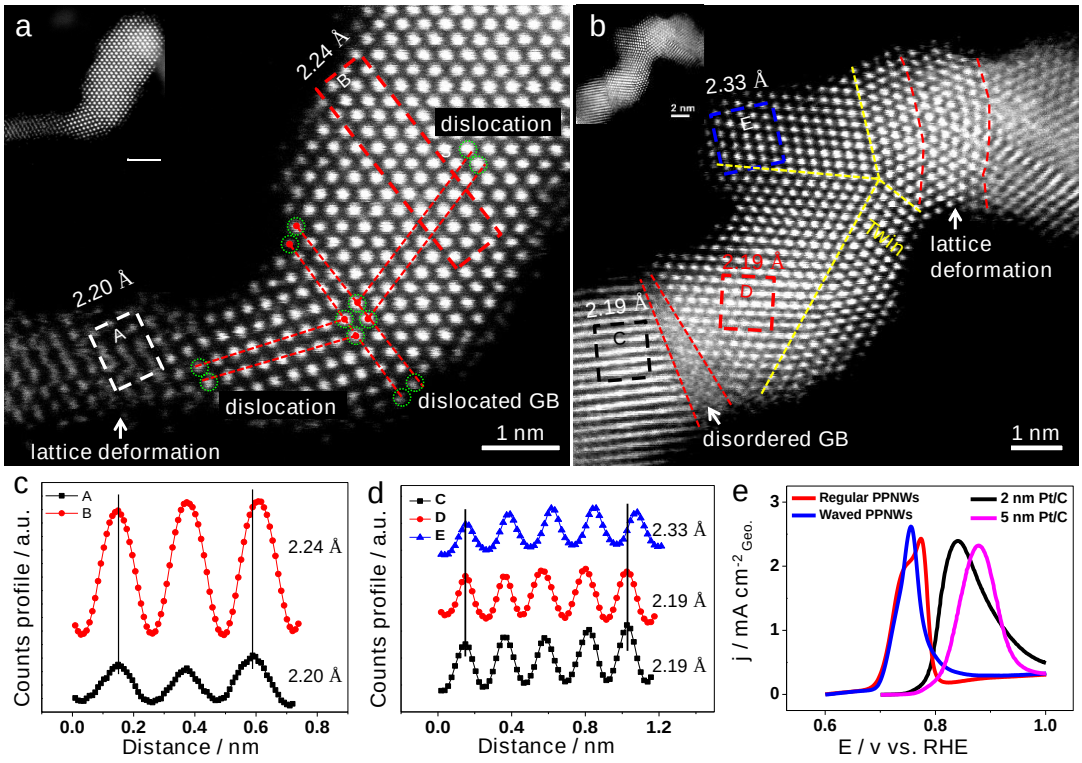

Figure 4: Direct observation of local lattice deformation and dislocations. (a-b) Atomically-resolved STEM images. (c-d) Inter-planar spacings on selected local lattices. (e) Background-subtracted CO_{ads} stripping voltammograms; Pt loadings for 5 nm Pt/C is 17.8 μg_{Pt} cm^{-2}_{Geo} while for others are 12.0 μg_{Pt} cm^{-2}_{Geo}. The red dash lines in (a) indicate the dislocation defects, along which the atoms deviate from the "correct" positions in the crystal structure.

We measured atomic-resolution STEM images to directly present local lattice deformations and dislocations around GBs. At a glance, GBs in PPNWs bring about substantially structural heterogeneities regarding atomic dislocation and lattice distortion. The d spacings of local lattices around GBs (see domains A-E in Figure 4 a-d) vary on a case-by-case basis depending on local micro-environments such as the distances from GBs and edge, MOAs of GBs, and types of GBs (twin, low-angle and high-angle GBs). Of note, the domains A, C, and D appear quite contracted against the bulk Pt (111). Besides, the dislocation-induced microstrain is observable in domains nearby GBs (see the analysis marked by the "dislocation" in Figure 4a). Meanwhile, GBs are a kind of structural defects to accommodate

geometrically-misoriented neighboring grains (see Figures 3-4), which could induce microstrain in and around GBs. Together, GBs not only themselves are disordered but also induce strain influence on neighboring grains. Expectedly, GBs and associated lattice microstrains tune d-state of surface Pt atoms and their affinity to reaction intermediates.[41, 42] The STEM presents the definite visual images for the lattice deformation and strain, but they might be a species of the statistically-limited information, which is complemented by the ensemble-level electrochemical and XAFS data.

We first estimated samples-averaging defects from a structurally-sensitive *CO stripping reaction.[29, 34, 36] A 5.0 nm Pt/C (TEC10E50E-HT, 50.9 wt.% obtained after a high-temperature heating treatment) was additionally examined for comparison. The kinetics of *CO electrooxidation is known accelerated by crystalline defects like undercoordinated sites via a bifunctional mechanism.[31, 43, 44] A cathodically-shifted *CO stripping peak on 2.0 nm Pt/C against 5 nm Pt/C (see Figure 4e) demonstrates the defects active for *CO electrooxidation because 2.0 nm Pt/C has a significantly higher density of undercoordinated sites than 5.0 nm Pt/C.[25] However, both regular and wave PPNWs display enhanced *CO electrooxidation kinetics compared to 2 nm Pt/C, hard to rationalize from the view of the undercoordinated step or edge density. The PPNWs form through an oriented attachment of grains accompanied with disappearance of edges at grain connections. Thereby, a higher step or edge density would not appear in PPNWs than in a similar-size nanoparticulate counterpart. Assuming *CO electrooxidation kinetics were promoted by weakening *CO binding energy, the GBs-induced microstrain (e.g. the compressive lattice microstrain in some deformed domains, see Figure 4 a-b) may facilitate *CO electrooxidation, which was recently suggested by Chattot et al.[29, 34] According to peak potentials of *CO electrooxidation,[29] the defects-induced microstrain follows a decreasing trend: waved PPNWs > regular PPNWs > 2 nm Pt/C.

3.3 In-situ XAFS analysis of structural defects in PPNWs

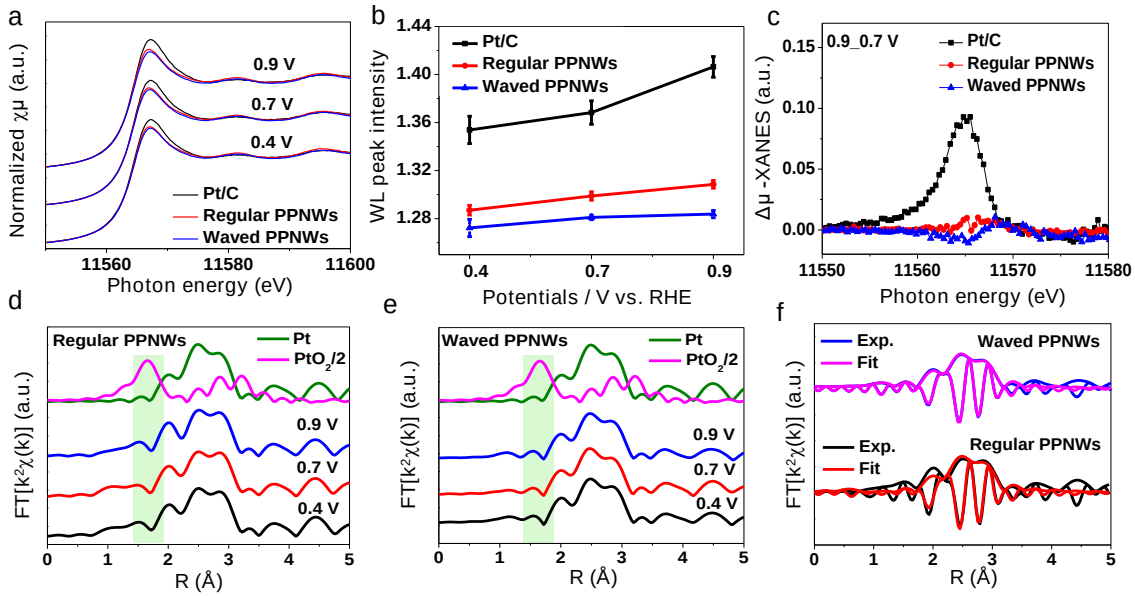

Figure 5: Operando XAFS to trace surface redox and structural parameters under ORR-relevant conditions. (a) XANES of working electrodes and (b) their white-line (WL) peak intensities. (c) Δμ-XANES spectra (0.9 V_0.7 V). (d-e) *In situ* EXAFS Fourier transforms of regular and waved PPNWs show only Pt-Pt scattering peaks when increasing potentials from 0.4 to 0.9 V_{RHE}. (f) Best fits of EXAFS for regular and waved PPNWs at 0.9 V_{RHE} in R space are obtained by only a Pt-Pt scattering shell.

The ORR is a structurally-sensitive reaction. Meanwhile, Pt in ORR experiences the surface redox, associated site-blocking effect and interfacial interaction with adsorbates. These interrelated surface and interfacial processes in turn affect the proceeding of ORR on Pt, making a distinction of structural effects considerably challenging.[45] To this end, we traced Pt surface under ORR-relevant potentials using *in situ* XAFS (Figure 5) in a potential-dependent way. Following by the ORR electrocatalysis, the *in situ* XAFS was measured (see experimental details in SI). For 2 nm Pt/C, the normalized white-line peak intensities of X-ray absorption near edge structure (XANES) at Pt-L_3 edge keep similar when increasing potentials from 0.4 to 0.7 V (vs. RHE) and then increases remarkably from 0.7 to 0.9 V due to the *OH/*O bonding and/or surface oxide generation on Pt surface. By contrast, white-line peak intensities remain almost unchanged from 0.4 to 0.9 V for both regular and waved PPNWs (Figure 5b). Both regular and waved PPNWs appear more metallic (i.e.

smaller white line intensity) than 2 nm Pt/C (Figure S4). We further used a $\Delta\mu_{Norm}$ method to reflect the amount and nature of adsorbates on electrodes (Figure 5c). The difference peak areas of $\Delta\mu_{Norm}$ (0.9_0.7V) = μ_{Norm}(0.9 V)-μ_{Norm}(0.7 V) decrease in the order: 2 nm Pt/C >> regular PPNWs ≈ waved PPNWs. Particularly, the difference peak intensities of $\Delta\mu_{Norm}$ (0.9_0.7V) for regular and waved PPNWs are one order of magnitude smaller than that for 2 nm Pt/C.[31, 46, 47] The *in situ* XANES spectra certify that *OH/*O species electrochemically adsorbed on Pt at 0.9 V are hard to detect by XANES for regular and waved PPNWs, which is entirely different from the case of 2 nm Pt/C. This aspect indicates that those *OH/*O species are quite disordered and thereby their contribution to XANES is weak and even negligible, for which plane-steps can account via their template roles to deform interfacial water networks and bring about a change in interfacial solvation effect relative to Pt (111).[31, 46-48] Consistent with XANES results, the best EXAFS fitting for 2 nm Pt/C needs a Pt–O scattering path due to the presence of discrete surface oxides (Figure S5 and Table S1). By contrast, the *in situ* EXAFS and their curve-fits suggest only a Pt-Pt scattering shell for regular and waved PPNWs (Figure 5d-f). Koper et al. recently demonstrated that adding Ni onto Pt(111) lowers barrier energies of interfacial water reorganization and thus facilitates proton transfer from water to electrodes[32]. Similarly, the proton transfer and oxygen gas transport from electrolyte to electrode in ORR require a significant rearrangement of interfacial water via hydrogen bonds[32, 49]. Given this, the disturbance of plane-steps to interfacial water via their template roles may accelerate proton transfer and consequently benefit proton-coupled electron transfer and ORR[7]. Notably, the edge sites of Pt NPs in 2 nm Pt/C did not result in remarkable disturbance of interfacial water networks, estimated from *in situ* XANES and EXAFS spectra in combination with electrochemical data (Figure 1i). Additionally, edge sites are undercoordinated and their binding to oxygen is too strong to directly electrocatalyze the ORR. Thus, the direct and/or indirect contribution from edge sites is small or negligible for the ORR.

As STEM images show different MOAs and associated microstrain between regular and waved PPNWs (Figure 2-4), we further discuss this aspect from the bulk-averaging or ensemble EXAFS. The best fits of EXAFS at 0.4 V (a potential at electric double layer) show that the averaged Pt-Pt interatomic distances (R_{Pt-Pt}) are 2.755±0.030, 2.755±0.009 and 2.758±0.007 Å for 2 nm Pt/C, regular PPNWs, and waved PPNWs, respectively (Table S1-S3). The difference in R_{Pt-Pt} within fitting error is quite small for the three catalysts and seems hard to interpret their remarkably distinct ORR and *CO electrooxidation activity. Note that the EXAFS-derived R_{Pt-Pt} is a bulk-averaging value probably masking the local heterogeneity, as shown in STEM images. We thus analyze their structural sources resulting in shortened R_{Pt-Pt} against bulk Pt (2.775 Å). The CN_{Pt-Pt} for 2 nm Pt/C is 8.3±2.1, which is smaller than 10.7± 2.4 and 11.1 ± 1.0, respectively for the regular PPNWs and waved PPNWs (Table S1-S3). Undercoordinated surface atoms contract inward[31, 50] either to stabilize the shared electron-pair bonds[51] or to smoothen the surface electronic density, [52, 53] interpreting a shortened R_{Pt-Pt} for 2 nm Pt/C with a high fraction of uncoordinated surface atoms, like edge and kink sites. However, a little larger CN_{Pt-Pt} for the regular and waved PPNWs is expected to result in a bit longer R_{Pt-Pt}, which cannot account for their quite similar R_{Pt-Pt} and thus suggest new structural sources to induce their shortened R_{Pt-Pt} rather than the mentioned coordination-dependent surface contraction.[31, 50] On this point, STEM images show direct evidence for the defects-induced microstrain (see Figure 4). This is also supported by the studies of Chattot et al.,[29, 34] who suggested that the structure-sensitive *CO electrooxidation can reflect the trend in lattice microstrain or surface lattice distortion: waved PPNWs > regular PPNWs > 2 nm Pt/C (see Figure 4e). The combined atomistic images, enhanced ORR and *CO electrooxidation activities, and *in situ* XANES and EXAFS spectra suggest that the plane-steps on stepped nanofacets and defect-induced microstrain promote the ORR for the regular and waved PPNWs. Given this, the unusual size effect for PPNWs could be rationalized. For the PPNWs with the same nanostructure, the smaller the grains are, the larger the GBs concentration becomes,

thus resulting in a larger fraction of strained surface atoms. In other words, the PPNWs with smaller diameters have more electroactive strained surface atoms that contribute to the enhancement of ORR activity.

4. Conclusion

We explored molecular-level electroactive factors in model PPNWs and revealed the promoting effect of GBs and plane-steps on the acidic ORR. As the local lattice strain increase with the MOAs of GBs, waved PPNWs with larger MOAs showed a better ORR activity than did regular PPNWs. The plane-steps disturbs interfacial water networks via preferable hydrogen bonding, which destabilizes ORR intermediates on terraces and presumably promotes proton transfer; while the edge-sites make almost no positive contribution to ORR activity despite the same coordination number between them. Our experimental results and insights presented here should be valuable for a spectrum of heterogeneous electrocatalytic materials.

Authorship contribution statement

Yasuhiro Iwasawa conceived the research and directed research. Xiao Zhao contributed to experiments and data analysis. Xiao Zhao and Shinobu Takao contributed to the HAADF-STEM analysis. Xiao Zhao, Yusuke Yoshida, Takuma Kaneko, Takao Gunji, Kotaro Higashi, and Tomoya Uruga contributed to the operando XAS analysis. The manuscript was primarily written by Xiao Zhao and Yasuhiro Iwasawa. All authors contributed to discussions and manuscript review.

Declaration of Competing Interest

The authors declare that they have no known competing financial interests or personal relationships that could have appeared to influence the work reported in this paper.

Data availability

Data will be made available on request.

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Appendix A. Supporting information

Supplementary data associated with this article can be found in the online version.

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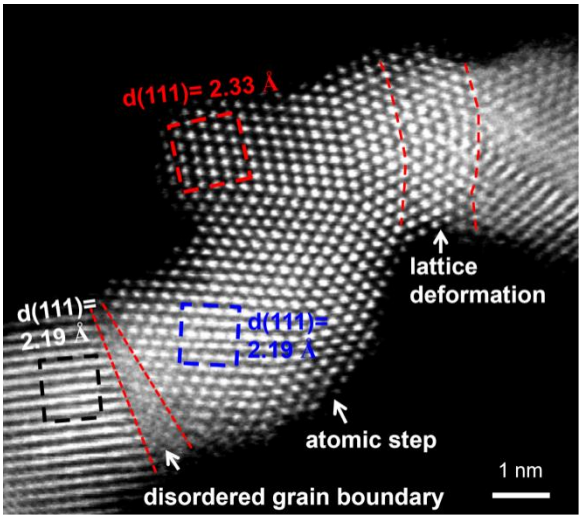
Graphical Abstract

Through synthesizing mode nanocatalysts, imaging atomistic structures and tracing dynamic operando spectra, we herein reveal the roles of structural defects in polycrystalline platinum nanowires for the enhanced oxygen electroreduction activity.



- The structural defects in polycrystalline platinum nanowires including grain boundaries and atomic edge-steps benefit oxygen reduction reaction.
- The grain boundaries induce local lattice strain that depends on misorientation angles.
- The edge-steps disturb interfacial water networks via preferable hydrogen bonding, which destabilizes oxygenated intermediates on terraces and presumably promotes proton transfer.

Declaration of interests

☒The authors declare that they have no known competing financial interests or personal relationships that could have appeared to influence the work reported in this paper.

☐The authors declare the following financial interests/personal relationships which may be considered as potential competing interests: